

MCQ of P-Block Elements

1. Which of the following represents the general electronic configuration of a p-block element?

- a) $(n-2)f^0 (n-1)d^0 ns^2 np^{1-6}$
- b) $(n-2)f^0 (n-1)d^{1-10} ns^2 np^{1-6}$
- c) $(n-2)f^1 (n-1)d^1 ns^2 np^{0-6}$
- d) $(n-2)f^0 (n-1)d^{1-10} ns^2 np^0$

Answer: a

2. Which group of the periodic table do p-block elements span?

- a) Group 1 to 2
- b) Group 3 to 12
- c) Group 13 to 18
- d) Group 17 and 18 only

Answer: c

3. As you move from left to right across a period in the p-block, the atomic size generally:

- a) Increases
- b) Decreases
- c) Remains constant
- d) First increases then decreases

Answer: b

4. Which of these is a noble gas (Group 18) p-block element?

- a) Oxygen
- b) Chlorine

Unit- 2 Compounds of p-Block Elements

c) Argon

d) Selenium

Answer: c

5. **Halogens belong to which group of the periodic table?**

a) Group 15

b) Group 16

c) Group 17

d) Group 18

Answer: c

6. **Which of the following is the most reactive halogen?**

a) Iodine

b) Bromine

c) Chlorine

d) Fluorine

Answer: d

7. **The atomicity of elemental phosphorus (white phosphorus) is:**

a) 2

b) 4

c) 6

d) 8

Answer: b

8. **Which of the following p-block elements has the highest electronegativity?**

a) Carbon

b) Nitrogen

c) Oxygen

Unit- 2 Compounds of p-Block Elements

d) Fluorine

Answer: d (fluorine is the most electronegative element overall)

9. **The ability of carbon to form long chains is called:**

a) Allotropy

b) Electronegativity

c) Catenation

d) Catalysis

Answer: c

10. **Which p-block group contains elements that are generally acidic in their oxide forms?**

a) Group 13

b) Group 14

c) Group 15

d) Groups 15–17 (non-metals)

Answer: d (Non-metal oxides tend to be acidic)